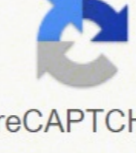


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## Neutralization Reactions Worksheet

Date \_\_\_\_\_

1. Write the balanced chemical equations for the neutralization reactions between:

- a) HI and NaOH  

$$\text{HI} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaI}$$
- b)  $\text{H}_2\text{CO}_3$  and  $\text{Sr}(\text{OH})_2$   

$$\text{H}_2\text{CO}_3 + \text{Sr}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{SrCO}_3$$
- c)  $\text{Ca}(\text{OH})_2$  and  $\text{H}_3\text{PO}_4$   

$$3 \text{Ca}(\text{OH})_2 + 2 \text{H}_3\text{PO}_4 \rightarrow 6 \text{H}_2\text{O} + \text{Ca}_3(\text{PO}_4)_2$$
- d) hydrobromic acid and barium hydroxide  

$$2 \text{HBr} + \text{Ba}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{BaBr}_2$$
- e) zinc hydroxide and nitric acid  

$$\text{Zn}(\text{OH})_2 + 2 \text{HNO}_3 \rightarrow 2 \text{H}_2\text{O} + \text{Zn}(\text{NO}_3)_2$$
- f) aluminum hydroxide and hydrochloric acid  

$$\text{Al}(\text{OH})_3 + 3 \text{HCl} \rightarrow 3 \text{H}_2\text{O} + \text{AlCl}_3$$

2. Complete and balance the following equations representing neutralization reactions:

- a)  $2 \text{CsOH} + \text{H}_2\text{CO}_3 \rightarrow 2 \text{H}_2\text{O} + \text{Cs}_2\text{CO}_3$
- b)  $2 \text{HF} + \text{Mg}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{MgF}_2$
- c)  $3 \text{HNO}_3 + \text{Al}(\text{OH})_3 \rightarrow 3 \text{H}_2\text{O} + \text{Al}(\text{NO}_3)_3$
- d)  $\text{HCl} + \text{KOH} \rightarrow \text{H}_2\text{O} + \text{KCl}$
- e)  $\text{HBrO}_3 + \text{LiOH} \rightarrow \text{H}_2\text{O} + \text{LiBrO}_3$

## Balancing Equations Race

- $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- $\text{Al} + \text{Fe}_3\text{N}_2 \rightarrow \text{AlN} + \text{Fe}$
- $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- $\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
- $\text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2$
- $\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{C}_7\text{H}_8 + \text{O}_2$
- $\text{NaClO}_3 \rightarrow \text{NaCl} + \text{O}_2$
- $(\text{NH}_4)_3\text{PO}_4 + \text{Pb}(\text{NO}_3)_4 \rightarrow \text{Pb}_3(\text{PO}_4)_4 + \text{NH}_4\text{NO}_3$
- $\text{BF}_3 + \text{Li}_2\text{SO}_3 \rightarrow \text{B}_2(\text{SO}_3)_3 + \text{LiF}$
- $\text{C}_7\text{H}_{17} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- $\text{CaCO}_3 + \text{H}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{CO}_3$
- $\text{Ag}_2\text{S} \rightarrow \text{Ag} + \text{S}_8$
- $\text{KBr} + \text{Fe}(\text{OH})_3 \rightarrow \text{KOH} + \text{FeBr}_3$
- $\text{KNO}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{K}_2\text{CO}_3 + \text{HNO}_3$
- $\text{Pb}(\text{OH})_4 + \text{Cu}_2\text{O} \rightarrow \text{PbO}_2 + \text{CuOH}$
- $\text{Cr}(\text{NO}_2)_2 + (\text{NH}_4)_2\text{SO}_4 \rightarrow \text{CrSO}_4 + \text{NH}_4\text{NO}_2$
- $\text{KOH} + \text{Co}_3(\text{PO}_4)_2 \rightarrow \text{K}_3\text{PO}_4 + \text{Co}(\text{OH})_2$
- $\text{Sn}(\text{NO}_2)_4 + \text{Pt}_3\text{N}_4 \rightarrow \text{Sn}_3\text{N}_4 + \text{Pt}(\text{NO}_2)_4$
- $\text{B}_2\text{Br}_6 + \text{HNO}_3 \rightarrow \text{B}(\text{NO}_3)_3 + \text{HBr}$
- $\text{ZnS} + \text{AlP} \rightarrow \text{Zn}_3\text{P}_2 + \text{Al}_2\text{S}_3$

1. The science of molecules and their transformations is known as \_\_\_\_\_.
2. Give several examples cited in the film where chemistry is involved in the real world.
3. The introduction of the lithium battery was important to what medical advance?
4. Why is it important to study the molecular structures and chemistry of living systems?
5. The mass of the materials coming out of a factory must be equal to what?
6. In addition to making new materials, chemists are also interested in what?
7. What famous chemist discovered the polypeptide structure known as the alpha helix?
8. Today's powerful microscopes are able to see individual \_\_\_\_\_.
9. What tool do scientists often use to study sub-microscopic structures?
10. What common compound is formed from sodium metal and chlorine gas?
11. What common compound is made by reacting salicylic acid with ethyl acetate?

Name : \_\_\_\_\_

Score : \_\_\_\_\_ Date: \_\_\_\_\_

## Net Ionic Equations

Balance the following equations and show their ionic & net ionic forms.



Ionic:

Net ionic:



Ionic:

Net ionic:



Ionic:

Net ionic:



Ionic:

Net ionic:



Ionic:

Net ionic:



Ionic:

Net ionic:



Ionic:

Net ionic:



Ionic:

Net ionic:

model problems explained step by step. There are so many awesome ways to teach toddlers... Composition of functions worksheet 1 answers 1. Which problem would you like to... 32-53) Acid and Base Hydrolysis VIDEO Hydrolysis of Cations VIDEO Hydrolysis of Anions VIDEO Hydrolysis of Salts VIDEO Hydrolysis of Amphoteric Anions VIDEO Introducing Metal and Non-metal Oxides VIDEO Metal Oxides VIDEO Non-metal Oxides VIDEO Oxides and Acid Rain VIDEO Balancing Neutralization Equations VIDEO Formula, Complete Ionic, and Net Ionic Equations for Neutralization VIDEO Acid-Base Titration Part 1  $i\frac{1}{2}$  Finding Unknown Concentration of an Acid VIDEO Acid-Base Titration Part 2  $i\frac{1}{2}$  Determining Percent Composition of Vinegar VIDEO Acid-Base Indicators  $i\frac{1}{2}$  How They Work VIDEO Using Indicators to Find Approximate pH VIDEO The Transition Point of an Indicator VIDEO DETAILED NOTES ON PRIMARY STANDARDS, pH OF ACID-BASE MIXTURES, TITRATION CURVES AND BUFFER SOLUTIONS (UNIT 4 NOTES P. 54-77) Strong Acid  $i\frac{1}{2}$  Strong Base Mixtures VIDEO Strong Acid-Strong Base Mixture Calculations Example 1 VIDEO Strong Acid  $i\frac{1}{2}$  Strong Base Mixture Calculations Example 2 VIDEO Strong Acid  $i\frac{1}{2}$  Strong Base Titration Curves VIDEO Strong Acid  $i\frac{1}{2}$  Weak Base Titration Curves VIDEO Weak Acid  $i\frac{1}{2}$  Strong Base Titration Curves VIDEO Buffer Solutions-Definition and Preparation VIDEO Buffer Solutions and How They Work  $i\frac{1}{2}$  Part 1 Adding an Acid VIDEO Buffer Solutions and How They Work  $i\frac{1}{2}$  Part 1 Adding a Base VIDEO Titration Simulation Site SOME REVIEW OF UNIT 4-2 (ACIDS, BASES AND SALTS) You can also choose to laminate them and use them with dry erase markers. 1 20 of 669 easy level free violin sheet music search within these results show levels. Letter Tracing Worksheets For Kindergarten Capital... She told herself that she would give it one more year and see how things went question 14. Print and download reflection sheet music from mulan arranged... B a ladder is leaning against the side of... More alphabet tracing free printable. Balancing Chemical Equations... Jon Feingersh Photography Inc / Getty Images Skip to the content Pre k kindergarten 1 st grade 2 nd grade 3 rd grade 4 th grade 5 th grade 6 th grade and 7 th grade.... Plus one page that contains all the letters upper and lower case at the same time. Coefficients equal to one 1 do not need to be shown in your answers. - Video DETAILED NOTES ON HYDROLYSIS, ACIDIC AND BASIC OXIDES (ANHYDRIDES), ACID RAIN, TITRATIONS AND INDICATORS (UNIT 4 NOTES P. Jose luis pelaez inc getty images good job.

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